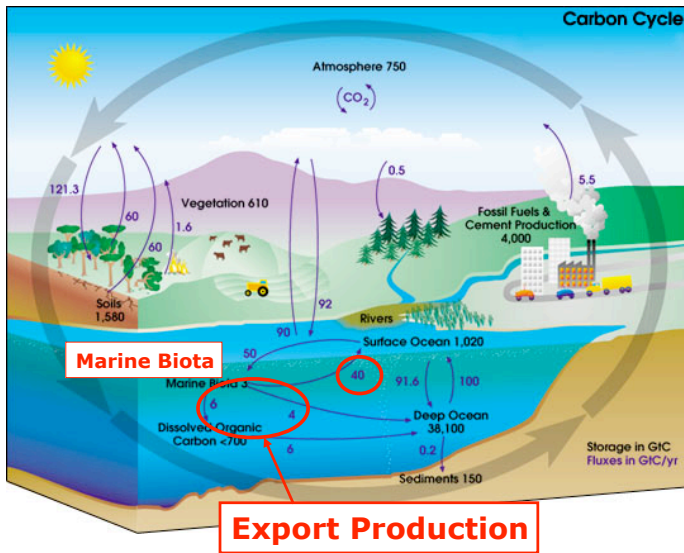
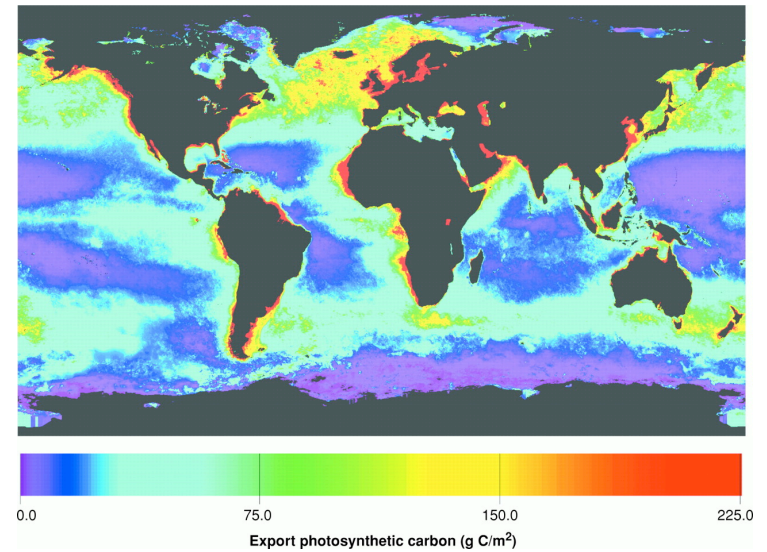


## Carbon Cycle

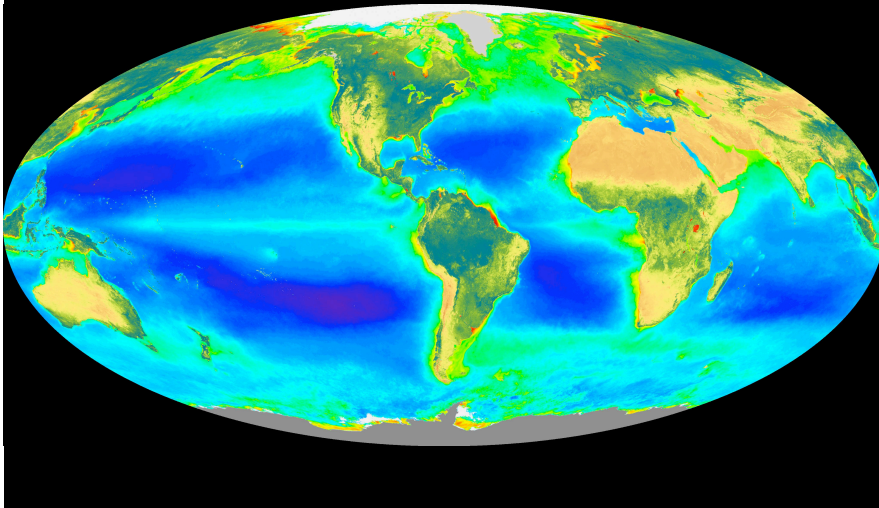


## Export Production of Organic Carbon



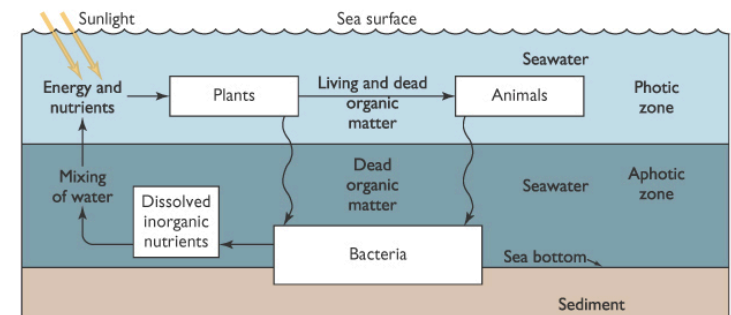
## Nitrogen and Iron cycles

Understanding the marine ecosystem dynamics



## Why do we care about Nitrogen?

- 1) N is a currency to track mass transfers in marine ecosystem
- 2) N dynamics important to model exchange processes



(a) NUTRIENT CYCLING

# Nitrogen

- N is an essential nutrient for all living organisms (nucleic acids and amino acids)
- N has many oxidation states, which makes speciation and redox chemistry very interesting
- $\text{NH}_4^+$  is preferred N nutrient

# Different forms of Marine N

Libes, 1992

### Bioavailable/Fixed (oxidation state)

$\text{NO}_3^-$   $5.7 \cdot 10^5$  Tg N (+5)

$\text{NO}_2^-$  500 Tg N (+3)

$\text{NH}_4^+$   $7.0 \cdot 10^3$  Tg N (-3)

Organic N  $5.3 \cdot 10^5$  Tg N (-3)

**Nitrate**

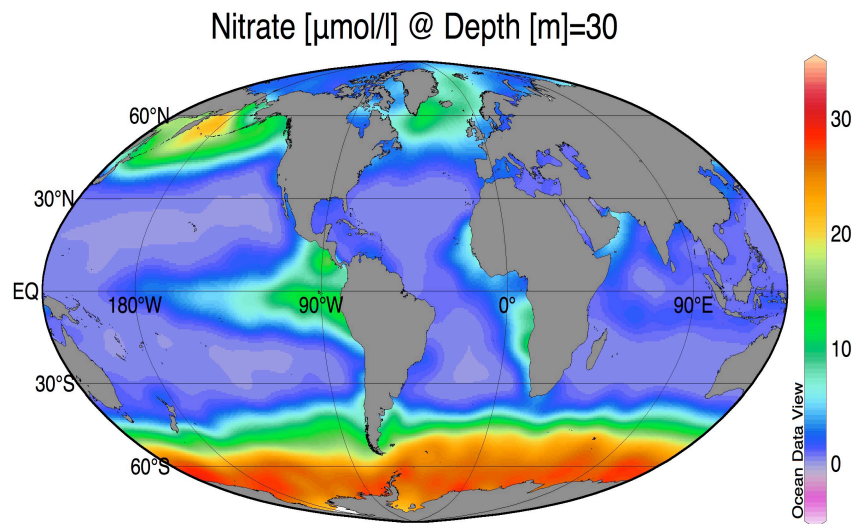
**Nitrite**

**Ammonia**

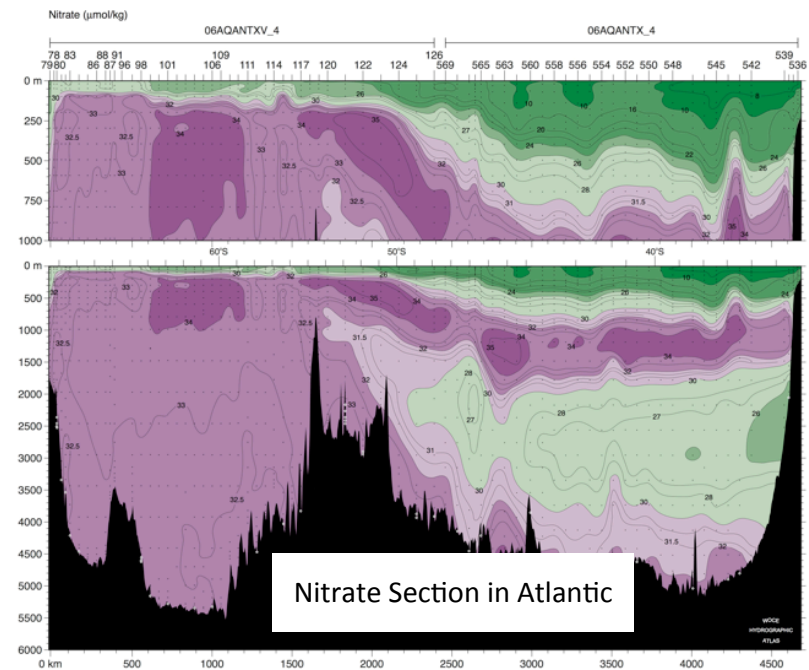
### Non-bioavailable

$\text{N}_2\text{O}$  200 Tg N (+1)

$\text{N}_2$   $2.2 \cdot 10^7$  Tg N (0)



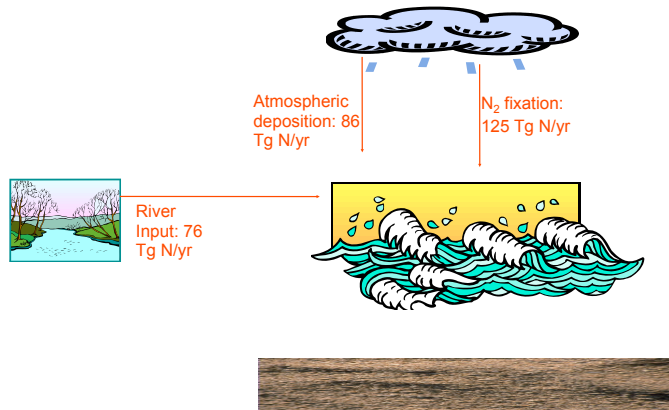
Data: eWOCE. Plot prepared with ODV



# Marine Fixed N Budget

Codispoti et al. (2001)

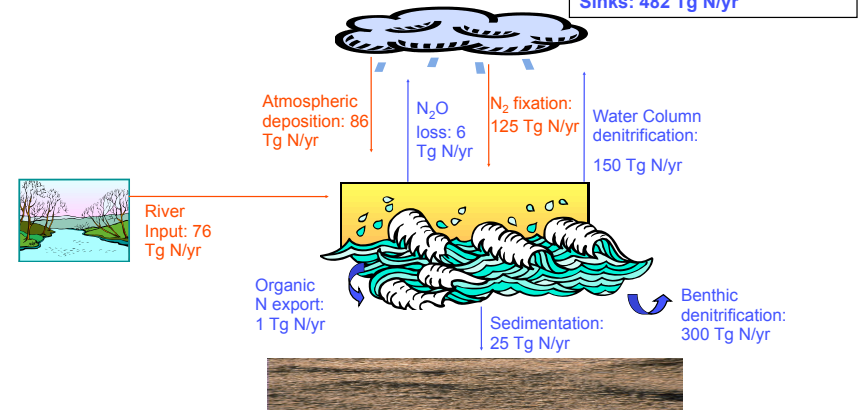
**Marine Reservoir:  $6.3 \times 10^5$  Tg N**  
Sources: 287 Tg N/yr



# Marine Fixed N Budget

Codispoti et al. (2001)

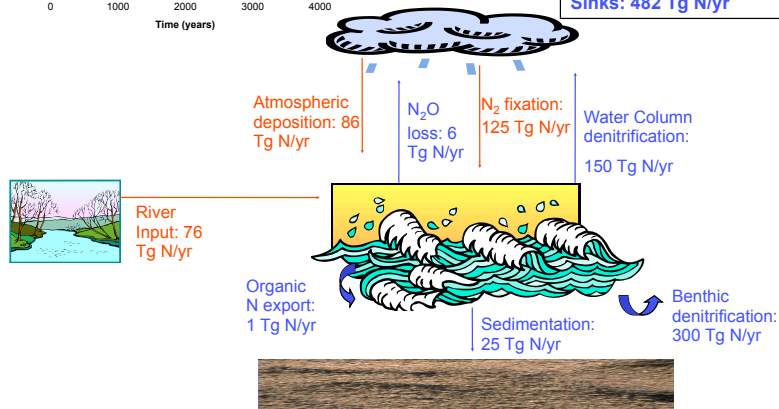
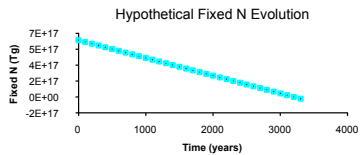
**Marine Reservoir:  $6.3 \times 10^5$  Tg N**  
Sources: 287 Tg N/yr  
Sinks: 482 Tg N/yr



# Marine Fixed N Budget

Codispoti et al. (2001)

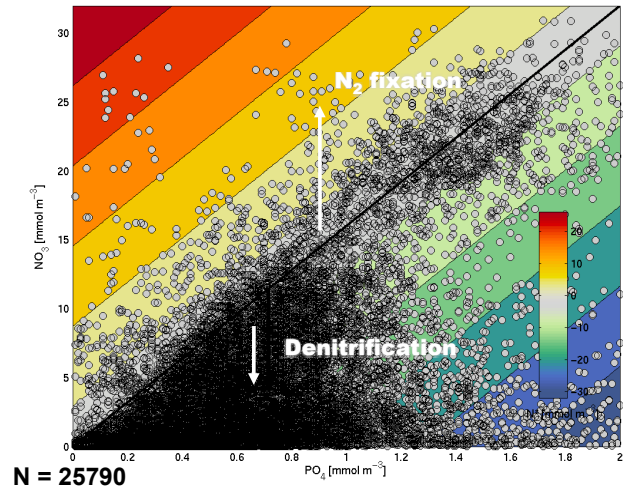
**Marine Reservoir:  $6.3 \times 10^5$  Tg N**  
Sources: 287 Tg N/yr  
Sinks: 482 Tg N/yr



# Oceanic Nitrogen Budget Estimates

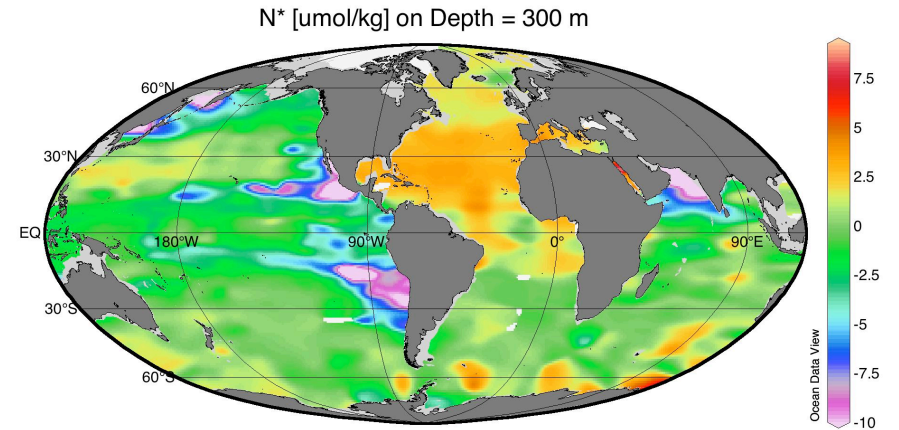
N Budget Terms (Tg N y <sup>-1</sup> )	1970	1979	1985	1997	2001
	(Delwiche)	(Liu)	(Codispoti & Christensen)	(Gruber & Sarmiento)	(Codispoti et al.)
<b>Inputs</b>					
atmospheric	4.1	49	40	15	56
runoff	30	17	25	41	41
N <sub>2</sub> -fixation	<b>10</b>	<b>30</b>	<b>25</b>	<b>125</b>	<b>125</b>
<b>Total Inputs</b>	<b>44.1</b>	<b>96</b>	<b>90</b>	<b>181</b>	<b>222</b>
<b>Outputs</b>					
pelagic denitrification	40	50	60	85	150
sedimentary denitrification	0	10	60	85	300
burial & other	0.2	36	38	19	32
<b>Total Outputs</b>	<b>40.2</b>	<b>96</b>	<b>158</b>	<b>189</b>	<b>482</b>

## Modern TIME



$N^* = N - 16 P$  (Gruber & Sarmiento 1997)

## N\* Distribution Shows Interplay Between N<sub>2</sub>-Fixation and Denitrification

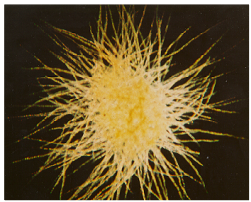


$N^* = 0.87([\text{NO}_3^-] - 16[\text{PO}_4^{3-}] + 2.9)$  (Gruber & Sarmiento 1997)

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## Trichodesmium: The Usual Suspect

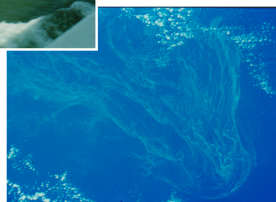
- Diazotrophs, including *Trichodesmium*, are broadly distributed in nutrient poor oceanic waters, but their contribution to the marine N budget remains poorly constrained.



*Trichodesmium* puffs (above) and tufts (right). Photos by Hans Paerl.

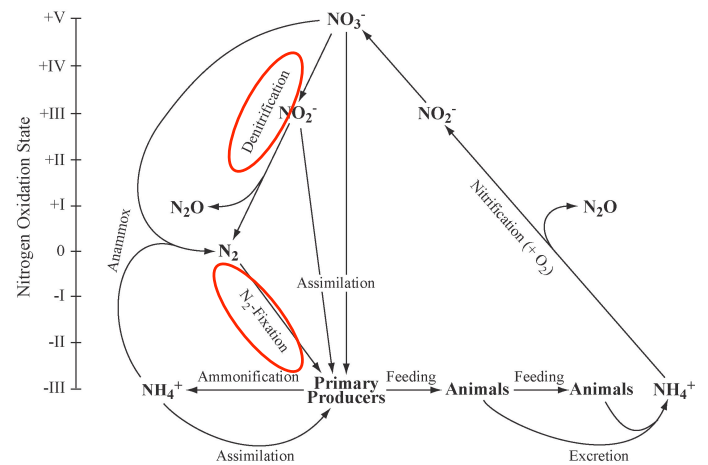


*Trichodesmium* blooms from aboard ship (left) and from space (below).



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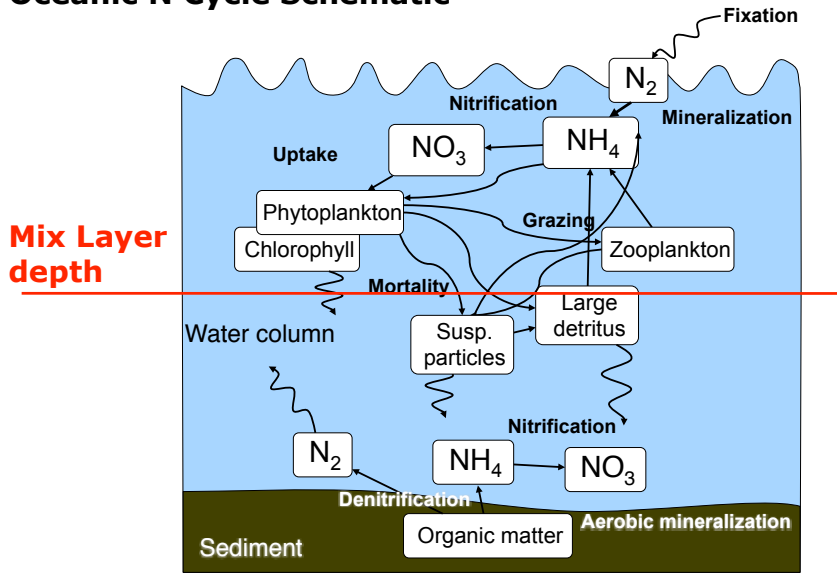
## Major Biological Transformations of Nitrogen



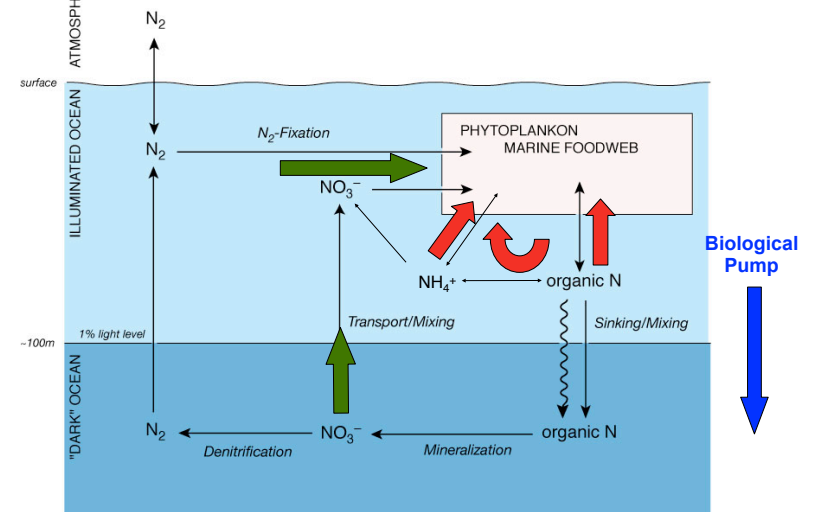
(Inspired by Codispoti 2001 and Liu 1979)

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## Oceanic N Cycle Schematic



## New vs. Regenerated Production

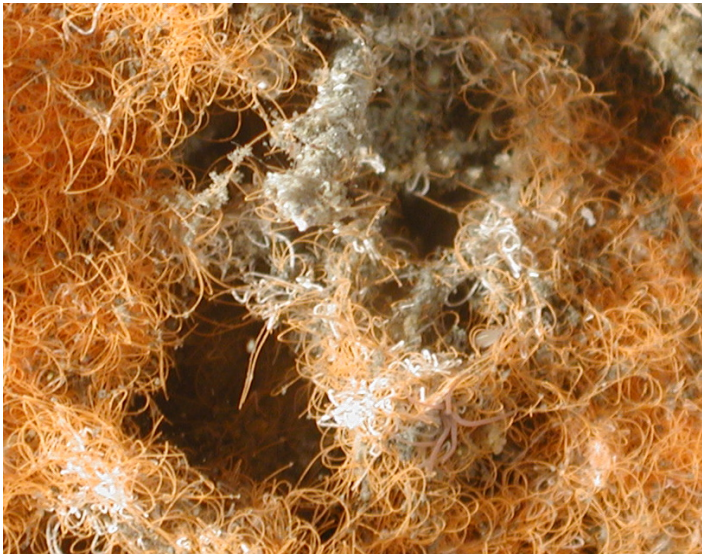


## Organic Matter Oxidation Sequence

Morel & Herring, 1993

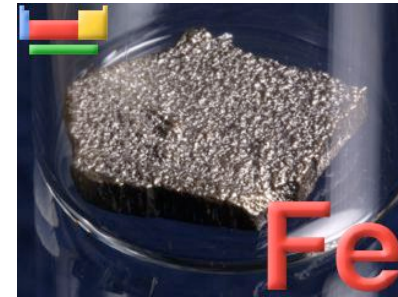
Process	$\Delta G^\circ$ (kJ/mol)
<b>Respiration</b> $\frac{1}{4}CH_2O + \frac{1}{4}O_2 \rightleftharpoons \frac{1}{4}CO_2 + \frac{1}{4}H_2O$	-119
<b>Denitrification</b> $\frac{1}{4}CH_2O + \frac{1}{5}NO_3^- + \frac{1}{5}H^+ \rightleftharpoons \frac{1}{4}CO_2 + \frac{1}{10}N_2 + \frac{1}{2}H_2O$	-113
<b>MnO<sub>2</sub> reduction</b> $\frac{1}{4}CH_2O + \frac{1}{8}MnO_2 + H^+ \rightleftharpoons \frac{1}{4}CO_2 + \frac{1}{2}Mn^{2+} + \frac{3}{4}H_2O$	-96.9
<b>Fe oxide reduction</b> $\frac{1}{4}CH_2O + Fe(OH)_3 + 2H^+ \rightleftharpoons Fe^{2+} + \frac{1}{4}CO_2 + \frac{11}{4}H_2O$	-46.7
<b>Sulfate reduction</b> $\frac{1}{4}CH_2O + \frac{1}{8}SO_4^{2-} + \frac{1}{8}H^+ \rightleftharpoons \frac{1}{8}HS^- + \frac{1}{4}CO_2 + \frac{1}{4}H_2O$	-20.5
<b>Methanogenesis</b> $\frac{1}{4}CH_2O \rightleftharpoons \frac{1}{8}CH_4 + \frac{1}{8}CO_2$	-17.7





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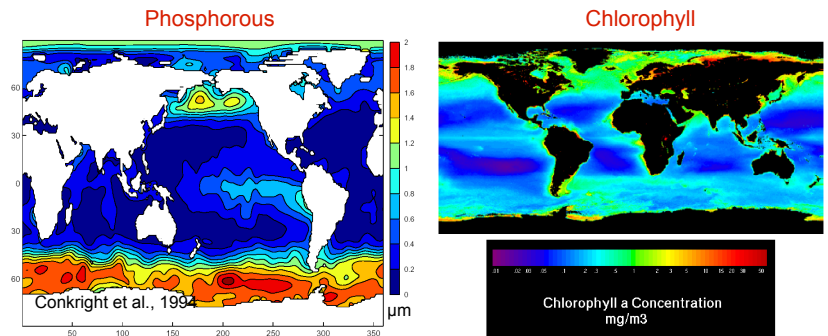
## Iron and Biogeochemical Cycles



### Redfield Ratio

- C:N:P
- 106:16 :1 (Redfield, 1958)
- **Could there be other essential micro-nutrients?**
- Trace metals such as Fe, Zn, Co are important!

### High Nutrient, Low (Medium) Chlorophyll Regions

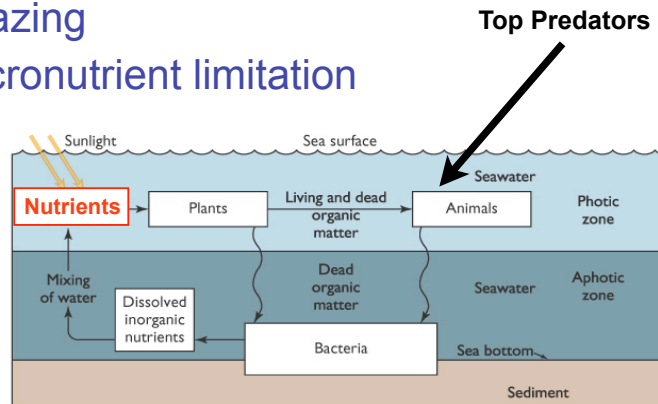


SeaWiFs

Why aren't the nutrients being completely utilized by phytoplankton?

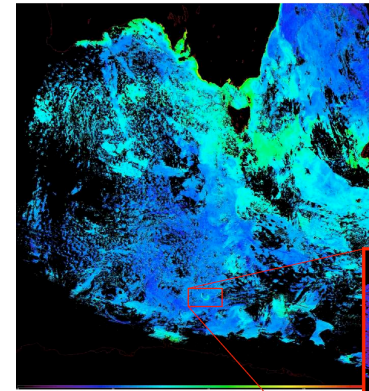
# Hypotheses

- Light
- Grazing
- Micronutrient limitation

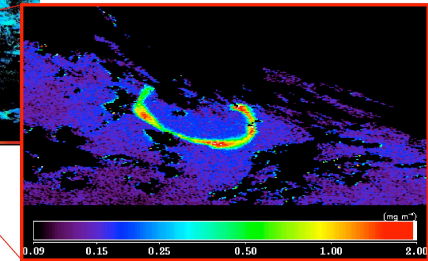


(a) NUTRIENT CYCLING

## In situ Fertilization experiments: Is iron limiting?

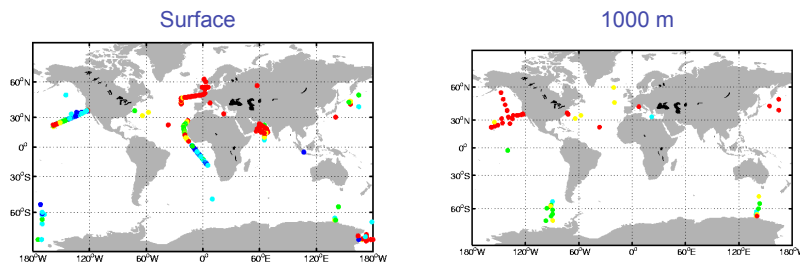


e.g. SOIREE  
Boyd et al., Nature (2000)



Iron needed for enzymes that facilitate electron transport, O<sub>2</sub> transport and other important functions.

## 'Dissolved' Iron distribution

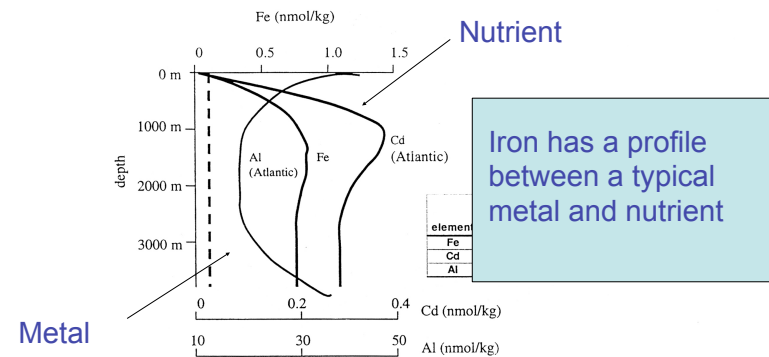


- >0.6 nM
- 0.4 – 0.6 nM
- 0.2 – 0.4 nM
- 0.1 – 0.2 nM
- <0.1 nM

Why are there so few measurements?

- Difficult to measure

## Iron Profile



Metal

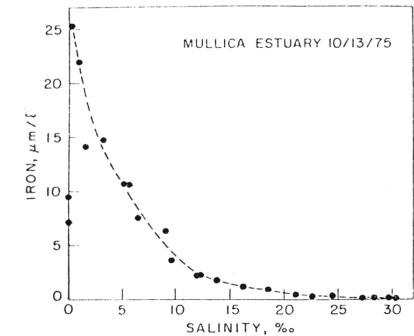
What controls the distribution (vertically and horizontally) of Iron?

## Sources of Iron

- Riverine
- Continental Shelves
- Dust

## Riverine

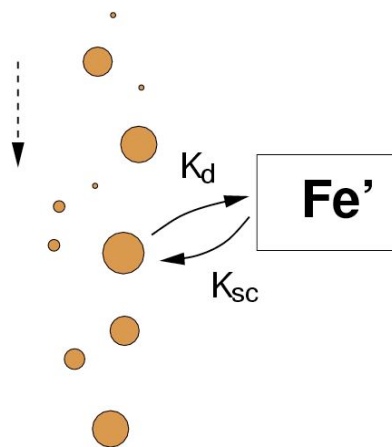
- $[Fe']$  decreases further from coast.
- This is due to scavenging of Fe by particles.
- We can conclude that rivers are not an important source for the open ocean



Boyle et al. (1977)

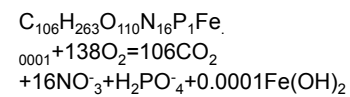
## Scavenging: Iron sink

- Iron lost to the ocean by scavenging – the process of sticking onto particles
- Rate of scavenging not well-known
- $loss = -k_{sc}[Fe'][P]$



## Continental Shelves

1. Resuspension of sediments can release Fe
2. When organic matter decomposes, Fe can diffuse or be bio-irrigated into the water column

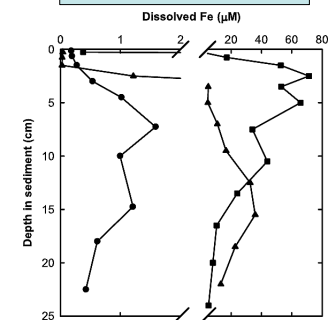


Estimate global flux of  $0.2-9 \times 10^{10} \text{ mol y}^{-1}$

Is this Fe upwelled to the surface before being scavenged?

Active area of research

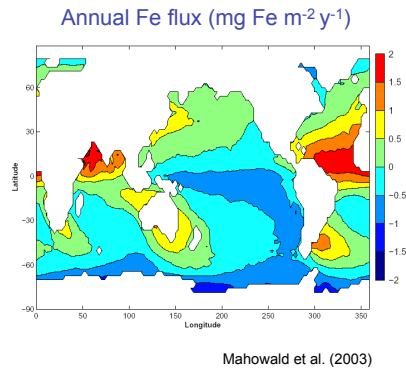
Results from flux chamber experiment (Elrod et al., 2004)



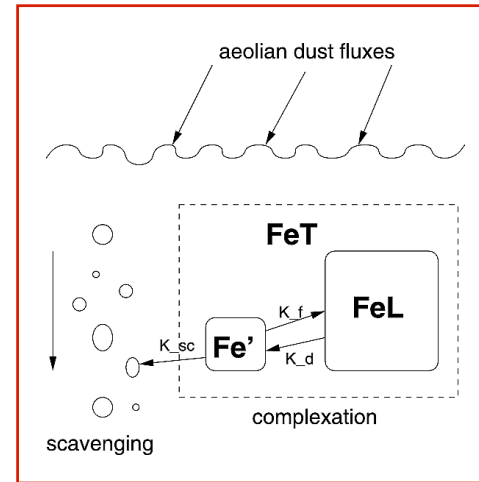


# Aeolian-derived Iron

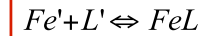
- Major source of iron
- How much of the iron is soluble?
  - 1-10%
- Active area of research: differences by provenance, processing in cloud, surface waters
- Flux:  $0.2-1.2 \times 10^{10} \text{ mol y}^{-1}$  (assuming 2% solubility)



# Iron Speciation : Complexation



$$Fe_T = Fe' + FeL$$



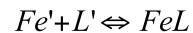
$$K = \frac{[FeL]}{[Fe'][L']}$$

$K$  = cond. stability constant  
specifies strength of ligand

# Iron Speciation : Complexation

- Inorganic iron:  $Fe^{2+}$ ,  $Fe^{3+}$ ,  $Fe(OH)_3$ 
  - Since ocean is oxidizing medium, reduced iron ( $Fe^{2+}$ ) concentrations are low.
  - Most  $Fe^{2+}$  produced by photochemistry, has a short lifetime
- 99% of Fe found bound to organic ligands
  - Increases solubility of iron in water column

$$Fe_T = Fe' + FeL$$

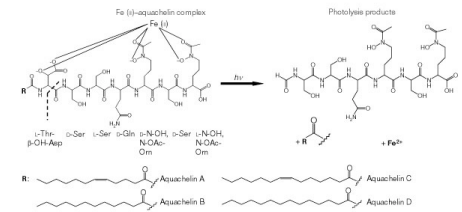


$$K = \frac{[FeL]}{[Fe'][L']}$$

$K$  = cond. stability constant  
specifies strength of ligand

# Complexation: Active areas of research

- What is the structure of the ligand?
  - messy organic molecular structure
- How do organisms produce it?
  - current research suggest marine bacteria produce the ligands.
- How do organisms utilize FeL?
  - Light breaks down FeL so organisms can grab the Fe'

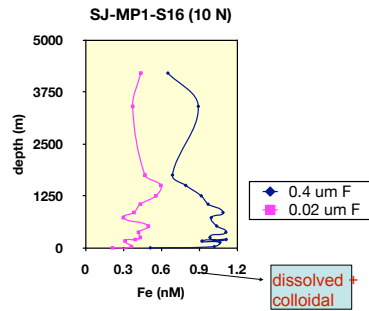


Barbeau et al. (2004)

# Forms of Iron

- Dissolved iron: <0.02  $\mu\text{m}$
- Colloidal: 0.02-0.4  $\mu\text{m}$
- Particulate: >0.4  $\mu\text{m}$

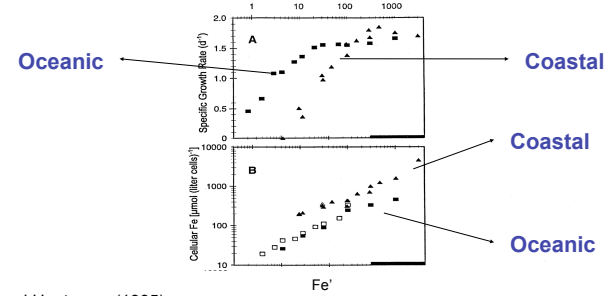
Active area of research: Role of colloidal matter



dissolved

Data from Boyle, 10N (Atlantic)

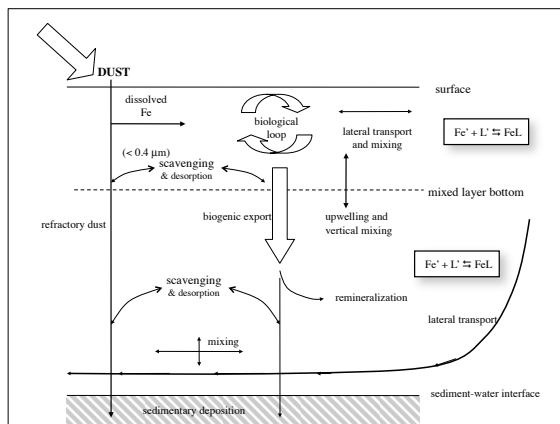
# Biological Uptake of Iron



Sunda and Huntsman (1995)

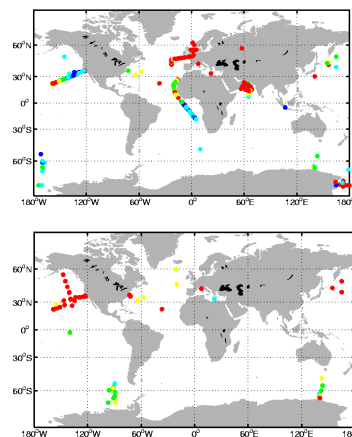
- Oceanic species have higher growth rates at lower  $[\text{Fe}]$
- They have adapted
  - Their Fe requirement is lower (small Fe:C ratio)
  - Oceanic species are smaller, so they have higher surface area:volume ratio

# Putting it all together



Developing mathematical model to understand the various processes affecting Fe

# Observations Model Results: Iron



- >0.6 nM
- 0.4 – 0.6 nM
- 0.2 – 0.4 nM
- 0.1 – 0.2 nM
- <0.1 nM

Parekh et al. (2004b)

## Link between dust flux and CO<sub>2</sub>?

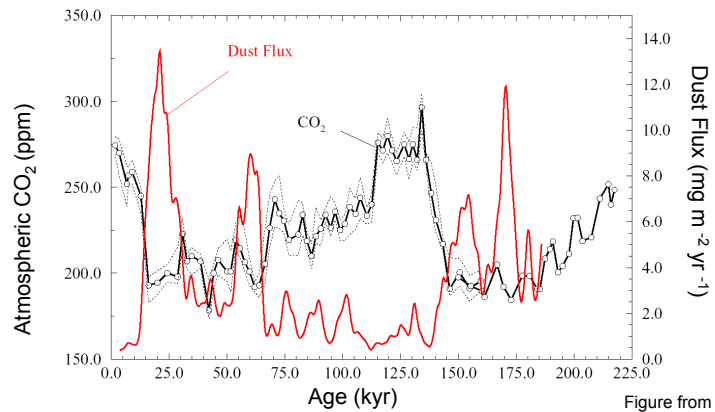
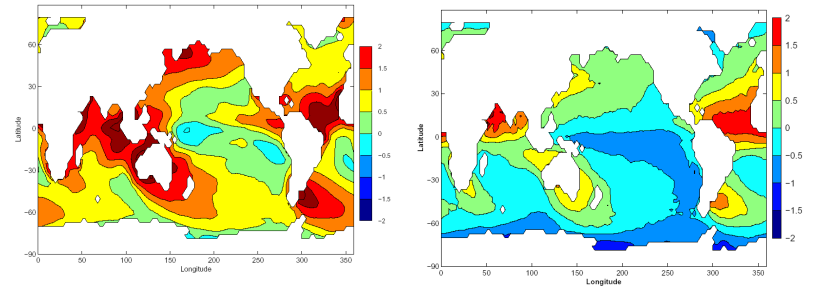


Figure from Gruber  
from Martin (1990)

+dust → +Fe → +bio. Productivity → +Export → +CO<sub>2</sub> drawdown

## Atmospheric CO<sub>2</sub> Sensitivity to Increased Dust Flux

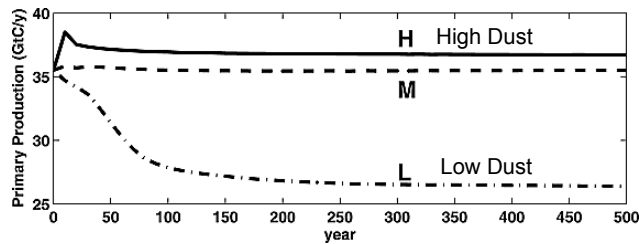


LGM dust flux

Present dust flux

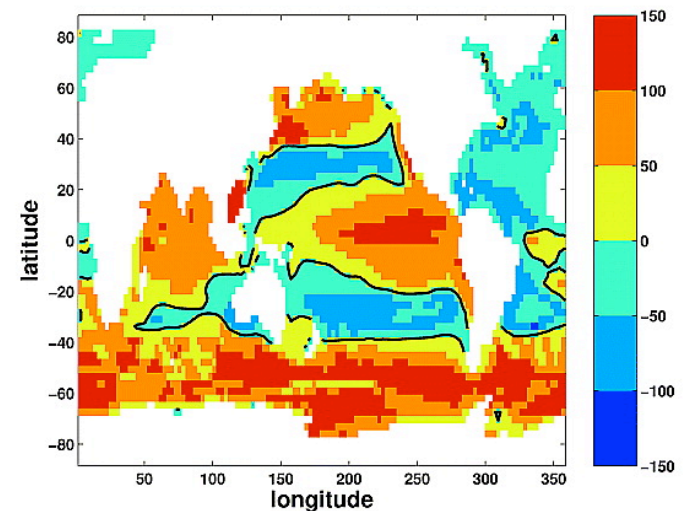
- 'Paleo' dust estimate from Mahowald *et al.* (1999)
- Dust flux greater 5.5 times globally

## Model result



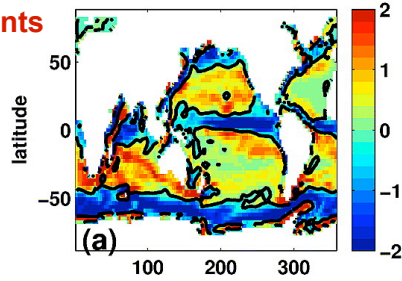
Time series of total global primary production (GtC yr<sup>-1</sup>) for high (solid line), medium (dashed line), and low (dash-dotted line) dust sensitivity studies.

**Difference in primary production (gC m<sup>-2</sup> yr<sup>-1</sup>)** between high and low dust sensitivity studies. Solid line is zero contour. Positive values indicate higher production when aeolian dust supply is enhanced.

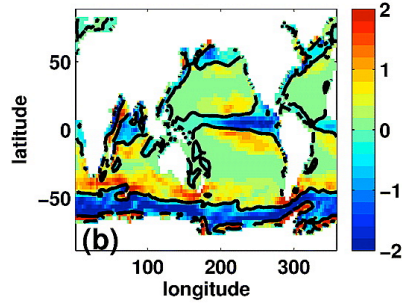


## Convergence of Macro Nutrients in surface waters

Low Dust

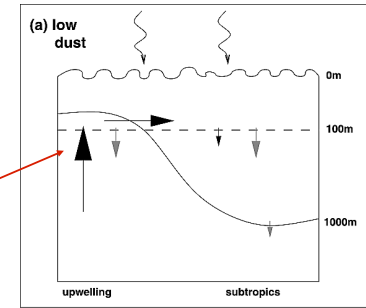


High Dust

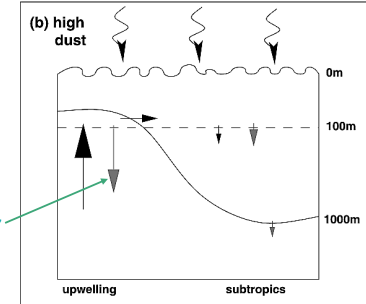


## Changes in Biogeochemical Cycling

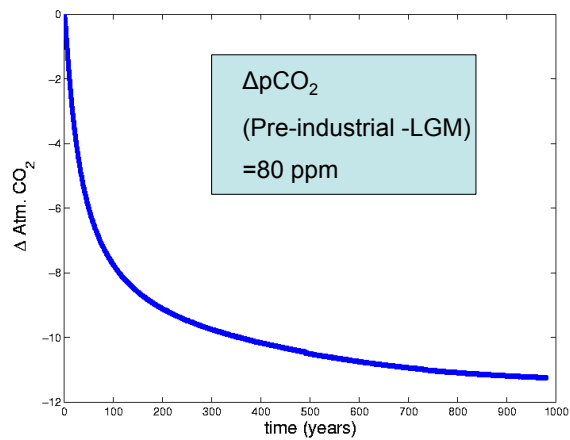
Macro Nutrients



Export of Organic Matter



## Model result



The effect of additional Fe is quite small.

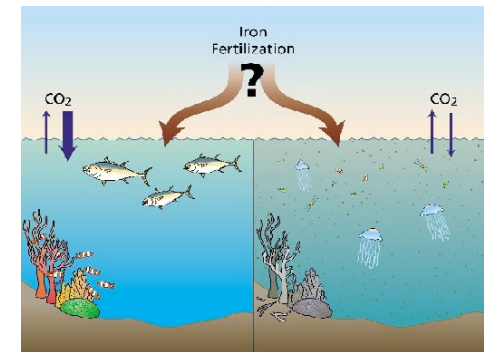
~11 ppm

## Iron Fertilization

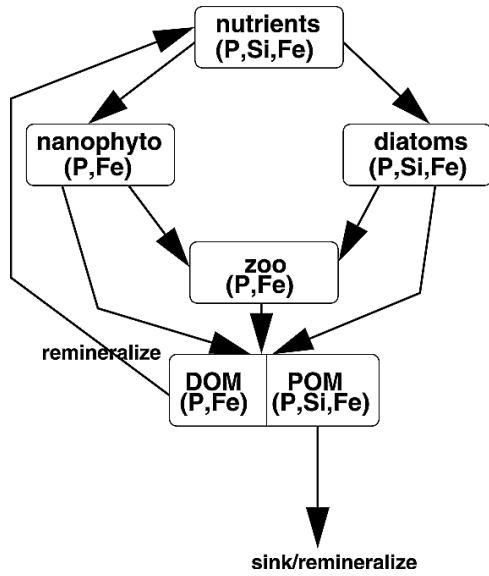
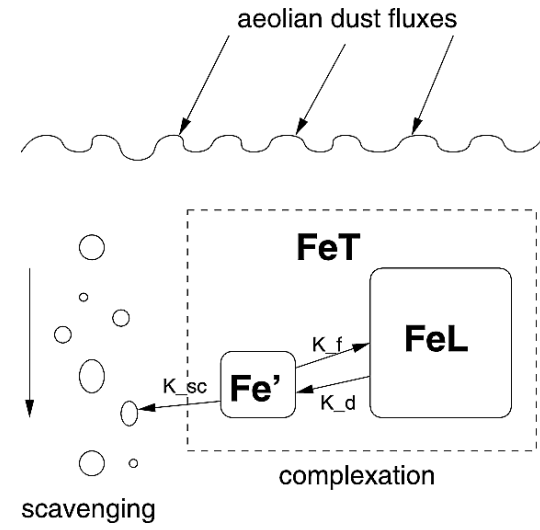
- Adding Fe artificially to transfer CO<sub>2</sub> from atmosphere to the sea

Open questions:

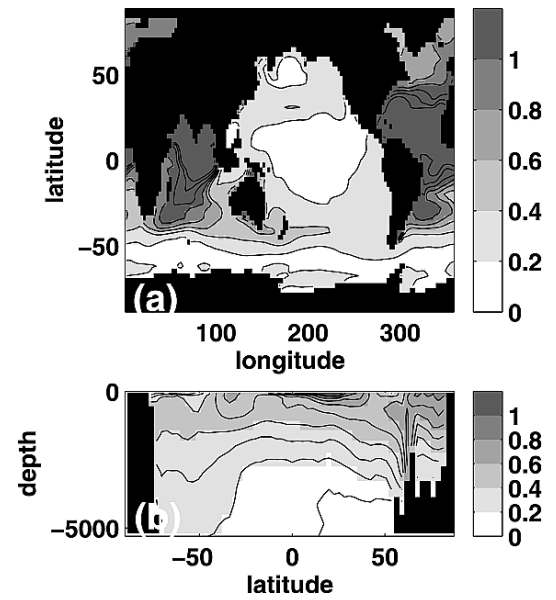
- How effective will it be?
- Effect on marine ecology?



End



Model IRON



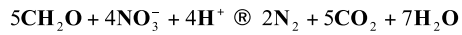
# Alternative pathways to N<sub>2</sub>

## Microbially mediated

### Nitrification



### Heterotrophic Denitrification

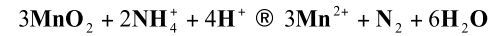


### Nitrogen Fixation

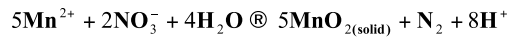


## Chemical Reactions

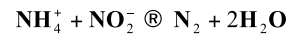
### MnO<sub>2</sub> Reduction



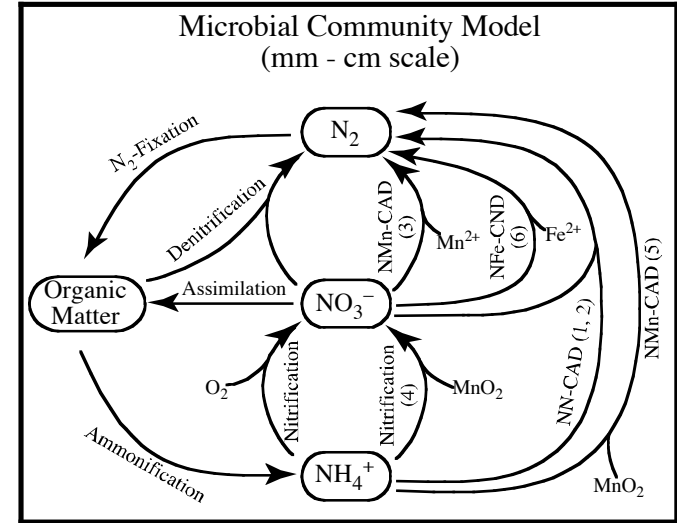
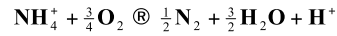
### Mn<sup>2+</sup> Oxidation



### Anammox

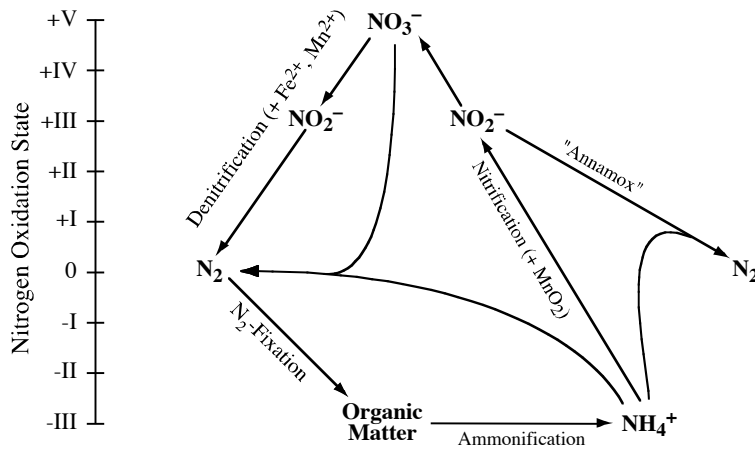


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# Alternative Pathways to N<sub>2</sub>



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